

# Example Of Law Of Definite Proportion

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A compound consists of 75.0g C and 12.0g of H. If I have a new sample of this compound that contains 35.2g C, how many grams of H do I have?

1. g H?

$$2. \quad \text{cpd} : \frac{75.0 \text{ g C}}{12.0 \text{ g H}} \quad \text{or} \quad \frac{12.0 \text{ g H}}{75.0 \text{ g C}}$$

new cpd: 35.2 g C

3. D.A.

$$(35.2 \text{ g C}) \left( \frac{12.0 \text{ g H}}{75.0 \text{ g C}} \right) = 5.632 \text{ g H}$$
$$= 5.63 \text{ g H}$$

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