

Example of work

Molecular Geometry Worksheet

Lewis Dot Structure (LDS) NH_3 <pre> xx H - N - H H </pre> Total # Ve^-: $1(5e^-) + 3(1e^-) = 8\text{ve}^-$ #/type of bonds: <u>3 single bonds</u> Central atom (Y/N) #LP: <u>1 LP</u> Geometry, Bond Angle: <u>Pyramidal</u> <u>107.0°</u> Hybridization: <u>sp^3</u>	Lewis Dot Structure (LDS) N_2 <pre> x x N ≡ N x x </pre> Total # Ve^-: $2(5e^-) = 10\text{ve}^-$ #/type of bonds: <u>1 double bond</u> Central atom (Y/N) #LP: <u>N/A</u> Geometry, Bond Angle: <u>Linear 180°</u> Hybridization: <u>none (no central atom)</u>
Lewis Dot Structure (LDS) Total # Ve^-: _____ #/type of bonds: _____ Central atom (Y/N) #LP: _____ Geometry, Bond Angle: _____ Hybridization: _____	Lewis Dot Structure (LDS) Total # Ve^-: _____ #/type of bonds: _____ Central atom (Y/N) #LP: _____ Geometry, Bond Angle: _____ Hybridization: _____

Note: Show all lone pair electrons (not just on central atom).