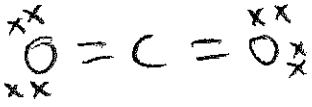
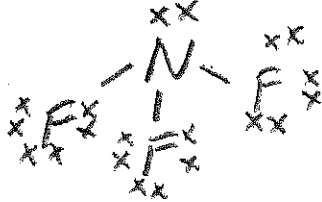


Example of work

Molecular Geometry Worksheet

Lewis Dot Structure (LDS) CO_2  Total # Ve^-: $1(4e^-) + 2(6e^-) = 16e^-$ #/type of bonds: <u>2 double bonds</u> Central atom (Y/N) #LP: <u>0 L.P.</u> Geometry, Bond Angle: <u>Linear 180°</u> Hybridization: <u>sp^2</u>	Lewis Dot Structure (LDS) NF_3  Total # Ve^-: $1(5e^-) + 3(7e^-) = 26e^-$ #/type of bonds: <u>3 single bonds</u> Central atom (Y/N) #LP: <u>1 L.P.</u> Geometry, Bond Angle: <u>Pyramidal 107.0°</u> (From VSEPR table) Hybridization: <u>sp^3</u>
Lewis Dot Structure (LDS) Total # Ve^-: _____ #/type of bonds: _____ Central atom (Y/N) #LP: _____ Geometry, Bond Angle: _____ Hybridization: _____	Lewis Dot Structure (LDS) Total # Ve^-: _____ #/type of bonds: _____ Central atom (Y/N) #LP: _____ Geometry, Bond Angle: _____ Hybridization: _____

Note: Show all lone pair electrons (not just on central atom).