Boyle's Law - One Problem

A gas occupies 2.45 L at a pressure of 475mmHg. If I decrease the pressure to 262mmHg, what is the new volume (in L) ?

$V_1 = 2.45\,l$

$P_1 = 475\,mmHg$

$P_2 = 262\,mmHg$

$V_2 = ?$

$V_2 = V_1\left(\frac{P_1}{P_2}\right)$

$V_2 = (2.45\,l)\left(\frac{475\,mmHg}{262\,mmHg}\right)$

$V_2 = 4.441\,l$

$\boxed{V_2 = 4.44\,l}$