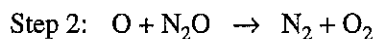
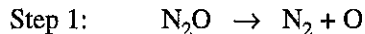


6. The table below gives data for a reaction rate study of the reaction represented above.

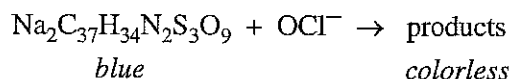
Experiment	Initial $[\text{H}_2]$ (mol L^{-1})	Initial $[\text{Cl}_2]$ (mol L^{-1})	Initial Rate of Formation of HCl ($\text{mol L}^{-1} \text{s}^{-1}$)
1	0.00100	0.000500	1.82×10^{-12}
2	0.00200	0.000500	3.64×10^{-12}
3	0.00200	0.000250	1.82×10^{-12}

- Determine the order of the reaction with respect to H_2 and justify your answer.
- Determine the order of the reaction with respect to Cl_2 and justify your answer.
- Write the overall rate law for the reaction.
- Write the units of the rate constant.
- Predict the initial rate of the reaction if the initial concentration of H_2 is $0.00300 \text{ mol L}^{-1}$ and the initial concentration of Cl_2 is $0.000500 \text{ mol L}^{-1}$.

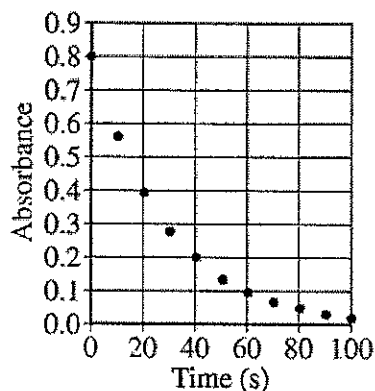
The gas-phase decomposition of nitrous oxide has the following two-step mechanism.



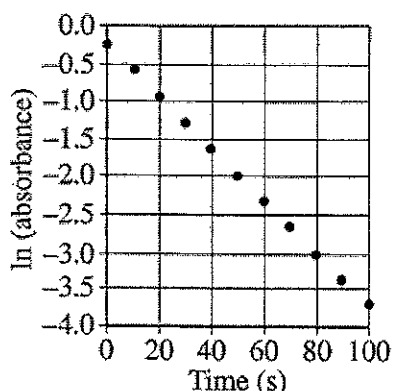
- Write the balanced equation for the overall reaction.
- Is the oxygen atom, O, a catalyst for the reaction or is it an intermediate? Explain.
- Identify the slower step in the mechanism if the rate law for the reaction was determined to be $\text{rate} = k [\text{N}_2\text{O}]$. Justify your answer.



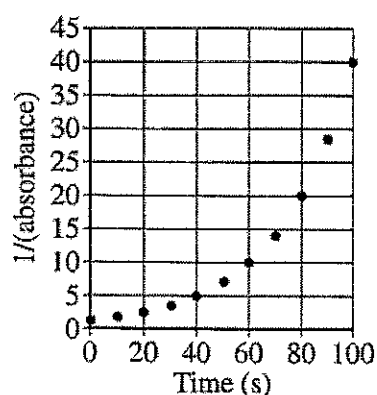
5. Blue food coloring can be oxidized by household bleach (which contains OCl^-) to form colorless products, as represented by the equation above. A student used a spectrophotometer set at a wavelength of 635 nm to study the absorbance of the food coloring over time during the bleaching process. In the study, bleach is present in large excess so that the concentration of OCl^- is essentially constant throughout the reaction. The student used data from the study to generate the graphs below.



Graph I



Graph II



Graph III

- (a) Based on the graphs above, what is the order of the reaction with respect to the blue food coloring?
- (b) The reaction is known to be first order with respect to bleach. In a second experiment, the student prepares solutions of food coloring and bleach with concentrations that differ from those used in the first experiment. When the solutions are combined, the student observes that the reaction mixture reaches an absorbance near zero too rapidly. In order to correct the problem, the student proposes the following three possible modifications to the experiment.

- Increasing the temperature
- Increasing the concentration of the food coloring
- Increasing the concentration of the bleach

Circle the one proposed modification above that could correct the problem, and explain how that modification increases the time for the reaction mixture to reach an absorbance near zero.

- (c) In another experiment, a student wishes to study the oxidation of red food coloring with bleach. How would the student need to modify the original experimental procedure to determine the order of the reaction with respect to the red food coloring?