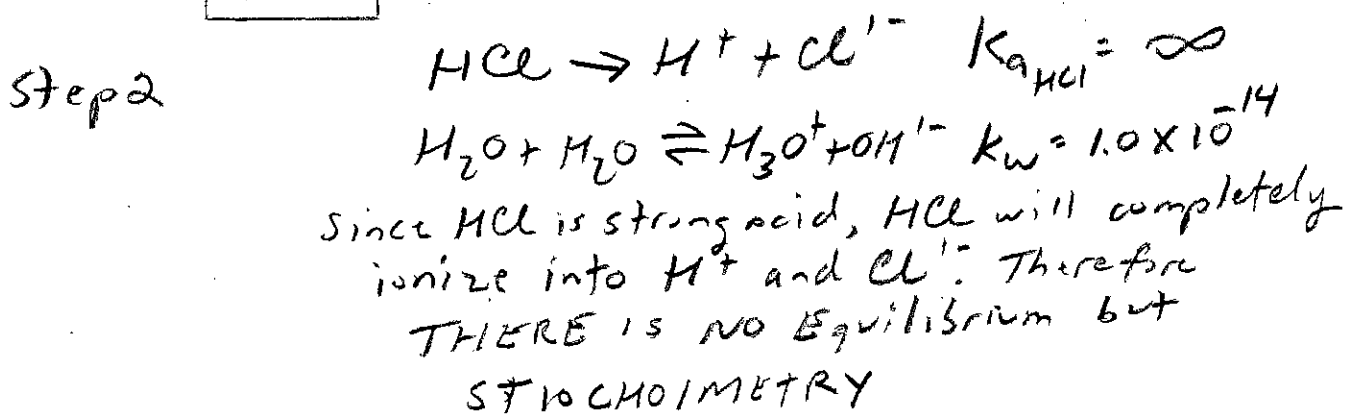
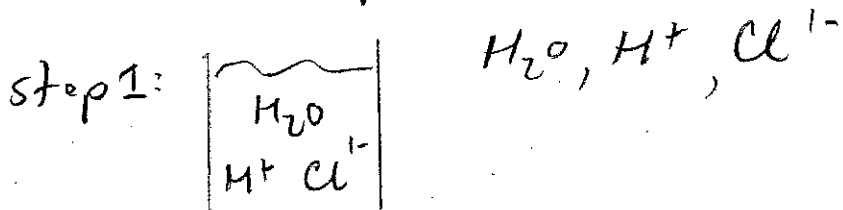


# pH Calculation of Strong Acid Solution

What is pH of a 0.10M HCl solution?



$$[\text{H}^+] = (0.10 \text{ M HCl}) \left( \frac{1 \text{ mole H}^+}{1 \text{ mole HCl}} \right) = 0.10 \text{ M}$$

step 6  $\text{pH} = -\log[\text{H}^+] = -\log(0.10 \text{ M}) = -(-1.000) = 1.00$

$\text{pH} = 1.00$